

SCIENCE

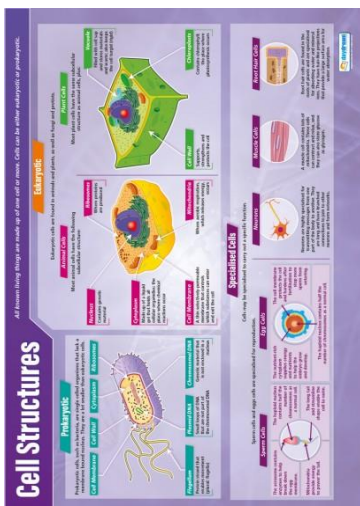
Posters

The posters are colourful and informative. Using the posters in the classroom, noticeboards, science labs and corridors makes learning fun and interesting and teaching becomes easy and effective too. They can also be used during events. E.g Science Week, etc.

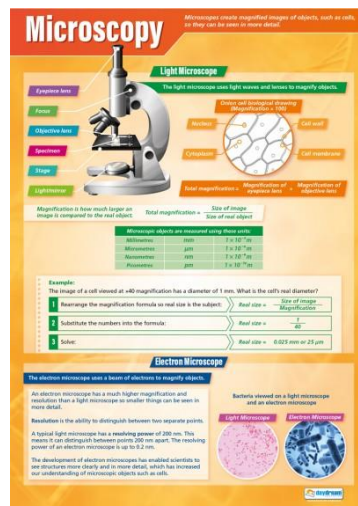
Reasons for using posters

1. About one third of students in an average classroom are visual learners.
2. Visual learners respond well to **colour**.
3. Images, photographs and diagrams are helpful learning aids for visual learners.
4. Words linked to pictures help visual learners grasp and remember new concepts.
5. Posters help reinforce important concepts and can be referred to regularly.
6. Posters can act as reference for students instead of asking the teachers.
7. Posters can keep your classroom/school fresh and **stimulating**.

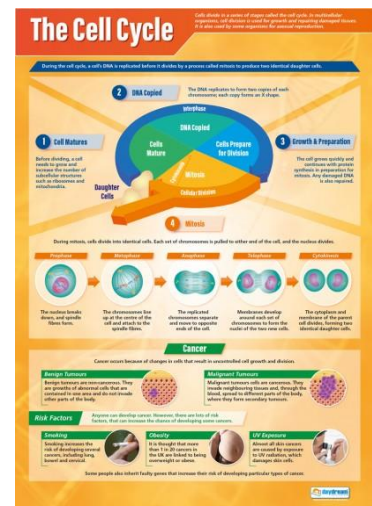
Posters – Biology



Cell Structures (A1 size)
Code: SC 59



Microscopy (A1 size)
Code: SC 60



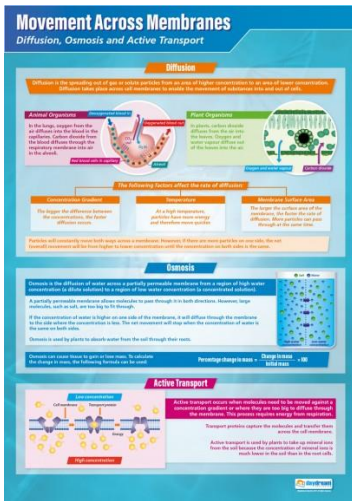
The Cell Cycle (A1 size)
Code: SC 61

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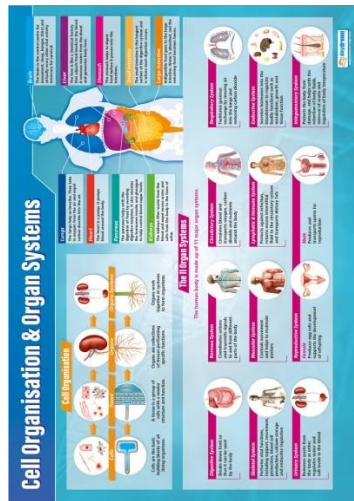
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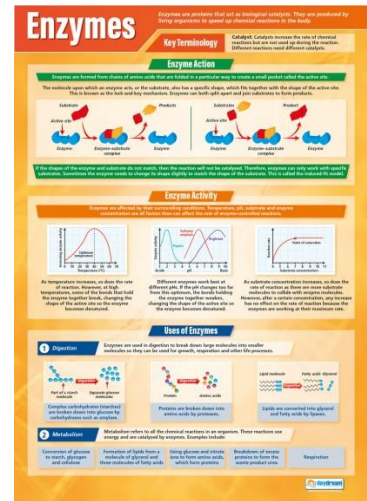
Posters – Biology



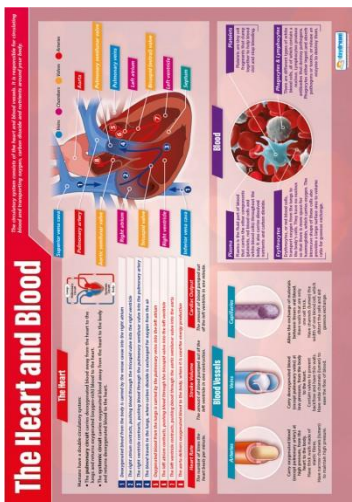
Movement Across Membranes
Code: SC 62 (A1 size)



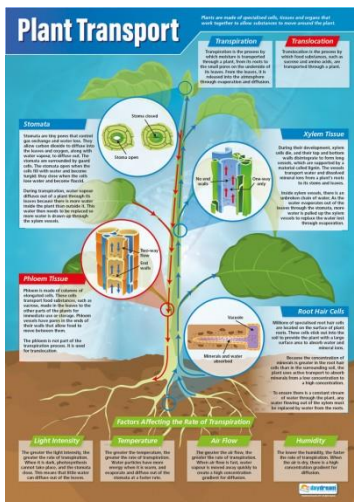
Cell Organisation & Organ Systems
(A1 size)
Code: SC 63



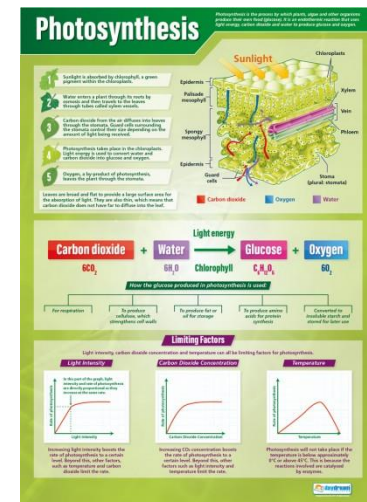
Enzymes (A1 size)
Code: SC 64



The Heart and Blood (A1 size)
Code: SC 65



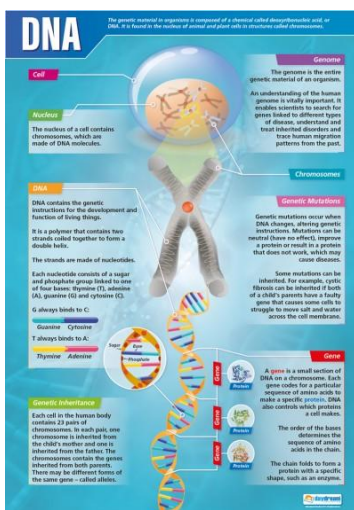
Plant Transport (A1 size)
Code: SC 66



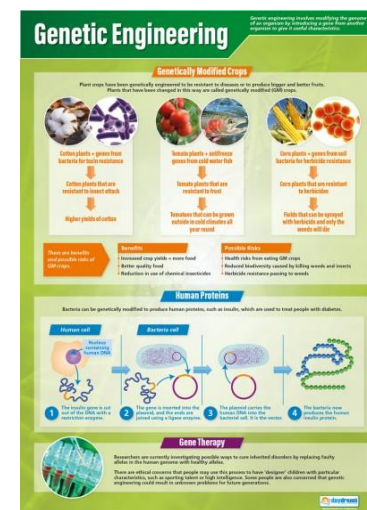
Photosynthesis (A1 size)
Code: SC 67



Fighting Disease (A1 size)
Code: SC 68



DNA (A1 size)
Code: SC 69



Genetic Engineering (A1 size)
Code: SC 70

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Posters – Biology and Chemistry

Classification of Living Organisms

Linnaeus's Classification Model

Traditionally, living things have been classified into groups based on their characteristics. This grouping system is called the classification system. The classification system is based on the characteristics of organisms. The classification system is based on the characteristics of organisms. The classification system is based on the characteristics of organisms.

Kingdoms: Animalia, Plantae, Fungi, Protista, Eubacteria, Monera.

Phyla: Chordata, Mollusca, Arthropoda, Annelida, Nematoda, Cnidaria, Porifera, Echinodermata, Gymnosperms, Angiosperms, Insecta, Mammalia, Reptalia, Amphibia, Fishes, Birds, Mollusca, Arthropoda, Annelida, Nematoda, Cnidaria, Porifera, Echinodermata, Gymnosperms, Angiosperms, Insecta, Mammalia, Reptalia, Amphibia, Fishes, Birds.

Three-Domain System

Improvements in technology and the understanding of molecular processes have meant that classification systems have had to be updated. Based on evidence from molecular biology, a three-domain system was proposed by Carl Woese in 1990.

Evolutionary Trees

Evolutionary trees are used to show the relationships between different groups of organisms. They show how groups of organisms have evolved from a common ancestor.

Classification of Living Organisms
Code: SC 71 (A1 size)

The Atom

All substances are made of atoms. An atom is the smallest part of an element that can exist.

Development of Atomic Theory

1803: Dalton's Atomic Theory. 1909: Rutherford's Atomic Model. 1913: Bohr's Atomic Model. 1926: Quantum Mechanical Model.

Atomic Structure of Carbon-12

Carbon-12 has 6 protons and 6 neutrons in its nucleus. It has 6 electrons arranged in two shells. The outer shell has 4 electrons.

Atomic Number & Mass Number

Atomic number is the number of protons in an atom. Mass number is the sum of protons and neutrons in an atom.

The Atom (A1 size)
Code: SC 72

Compounds

Compounds contain two or more elements that are chemically combined.

Chemical reactions always involve the formation of one or more new substances and other involve a change.

Chemical Equations

Chemical equations are used to represent chemical reactions. They show the reactants and products of a reaction.

Chemical Equations

Chemical equations are used to represent chemical reactions. They show the reactants and products of a reaction.

Compounds (A1 size)
Code: SC 73

Chemical Equations

A chemical equation is a symbolic representation of a chemical reaction. It shows the reactants and products of a reaction.

Writing Chemical Equations

Step 1: Write the word equation. Step 2: Write the symbols for the reactants and products. Step 3: Balance the equation. Step 4: Add state symbols.

Examples:

Hydrogen + Oxygen → Water
 $2H_2 + O_2 \rightarrow 2H_2O$

Chemical Equations (A1 size)
Code: SC 74

Separating Mixtures

Mixtures can be separated by physical processes. The type of process used depends on the substances in the mixture.

Key Terminology: Filtrate, Residue, Evaporate, Crystallise, Distillate, Residue, Fractional Distillation.

Separation Techniques: Filtration, Evaporation and Crystallisation, Distillation, Paper Chromatography, Fractional Distillation.

Separating Mixtures (A1 size)
Code: SC 75

Groups in the Periodic Table

The periodic table is divided into groups. Each group contains elements with similar properties.

Group 1: Alkali Metals

Group 2: Alkaline Earth Metals

Group 7: Halogens

Group 8: Noble Gases

Groups in the Periodic Table
Code: SC 76 (A1 size)

Bonding

Bonds are formed of attraction that hold atoms together. There are three types of bonding: ionic, covalent, and metallic.

Ionic Bonding

Ionic bonding occurs between a metal and a non-metal. The metal atom loses electrons to form a cation, and the non-metal atom gains electrons to form an anion.

Covalent Bonding

Covalent bonding occurs between two non-metals. The atoms share their outer electrons to form a stable arrangement.

Metallic Bonding

Metallic bonding occurs in metals. The metal atoms are arranged in a regular lattice, and the outer electrons are shared among all the atoms.

Bonding (A1 size)
Code: SC 77

Carbon

Carbon has several allotropes: diamond, graphite, and fullerenes.

Diamond: A hard, transparent crystal. Each carbon atom is bonded to four other carbon atoms in a tetrahedral arrangement.

Graphite: A soft, black, flaky substance. Each carbon atom is bonded to three other carbon atoms in a hexagonal arrangement.

Fullerenes: A form of carbon that has a spherical structure. Each carbon atom is bonded to three other carbon atoms.

Carbon (A1 size)
Code: SC 78

States of Matter

The three states of matter are solid, liquid, and gas.

Solid: Particles are packed closely together and vibrate in fixed positions.

Liquid: Particles are close together but can move past each other.

Gas: Particles are far apart and move rapidly in all directions.

State Changes: Melting, Freezing, Boiling, Condensation, Evaporation, Sublimation, Deposition.

States of Matter (A1 size)
Code: SC 79

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Posters – Chemistry & Physics

Acids and Alkalis

The pH Scale

The pH scale is a measure of the acidity or alkalinity of a solution. It is measured by universal indicators, which change colour depending on the pH of the solution. The colour for each pH is shown below. A table of acids and alkalis is also provided to help you identify the pH of a solution.

Acids (pH 0-6) and **Alkalis** (pH 8-14) are shown with their respective colors and strengths.

Reactions of Acids

- Acids react with metals to produce salts. The salt formed is named after the metal and the acid.
- Acids react with carbonates to produce salts, water, and carbon dioxide.
- Acids react with bases to produce salts and water.

Reactions of Alkalis

- Alkalis react with acids to produce salts and water.
- Alkalis react with metal ions to produce precipitates.

Acids and Alkalis (A1 size)
Code: SC 80

Reactivity of Metals

The Reactivity Series

The reactivity series is a list of metals arranged in order of their reactivity. The most reactive metal is at the top, and the least reactive is at the bottom.

Displacement Reactions

A more reactive metal will displace a less reactive metal from its compound.

Reactions with Acids and Water

Metals react with acids to produce hydrogen gas and a salt. Some metals also react with water to produce hydrogen gas and a metal hydroxide.

Metal Extraction

Metals are extracted from their ores using various methods, including reduction with carbon or electrolysis.

Reactivity of Metals (A1 size)
Code: SC 81

Electrolysis

Electrolysis of Molten Compounds

Electrolysis is a process that uses electricity to break down a compound into its constituent elements.

Electrolysis of Molten Lead(II) Bromide

At the cathode: $Pb^{2+} + 2e^{-} \rightarrow Pb$

At the anode: $2Br^{-} \rightarrow Br_2 + 2e^{-}$

Electrolysis of Aqueous Solutions

Electrolysis of aqueous solutions involves the discharge of ions from the solution.

Electrolysis to Extract Metals

Electrolysis is used to extract metals from their ores, such as aluminium and sodium.

Electrolysis (A1 size)
Code: SC 82

Exothermic & Endothermic Reactions

Exothermic Reactions

In an exothermic reaction, heat energy is given out, and the temperature of the surroundings increases.

Endothermic Reactions

In an endothermic reaction, heat energy is taken in, and the temperature of the surroundings decreases.

Reaction Profiles

Reaction profiles are used to show the relative energies of the reactants and products in a reaction, and the activation energy.

Measuring Energy Transfer

The amount of energy transferred in a chemical reaction is measured by measuring the temperature change using the following method:

1. Add one reactant to the cup and measure the temperature.
2. Add the other reactant and mix.
3. Measure the temperature of the solution at the end of the reaction.
4. If the temperature increases, the reaction is exothermic. If it decreases, the reaction is endothermic.

Exothermic & Endothermic Reactions (A1 size)
Code: SC 83

Crude Oil

Crude Oil Refining

Crude oil is refined into various products using a process called fractional distillation.

Cracking

Cracking is a process that breaks down large hydrocarbon molecules into smaller molecules.

Alkylates

Alkylates are used to improve the octane rating of petrol.

Asphaltenes

Asphaltenes are the heaviest fraction of crude oil, used in road building.

Crude Oil (A1 size)
Code: SC 84

Sustainable Development

The Six Rs of Sustainable Development

Reduce, Reuse, Recycle, Repair, Refuse, Rethink.

Recycling Materials

Recycling materials helps to conserve resources and reduce waste.

Life Cycle Assessment (LCA)

LCA is a process that evaluates the environmental impacts of a product throughout its life cycle.

Sustainable Development (A1 size)
Code: SC 85

Energy Stores & Systems

Energy Stores

Energy can be stored in various ways, such as in chemical stores, magnetic stores, and nuclear stores.

Energy Transfer

Energy is transferred between systems and objects in a system through heating, work done by forces, work done when a current flows, and radiation.

Calculating Energy Stores

The amount of energy in some of the most common energy stores can be calculated using the following formulas:

- Kinetic Energy Store: $E_k = \frac{1}{2}mv^2$
- Gravitational Potential Energy Store: $E_p = mgh$
- Elastic Potential Energy Store: $E_e = \frac{1}{2}kx^2$

Energy Stores & Systems (A1 size)
Code: SC 86

Energy Transfer and Efficiency

Energy Transfer

Energy can be transferred through conduction, convection, and radiation.

Conservation of Energy

Energy is conserved in a closed system, meaning it cannot be created or destroyed.

Efficiency

Efficiency is a measure of the proportion of the total energy input that is useful energy output.

$$\text{efficiency} = \frac{\text{useful output}}{\text{total input}} \times 100\%$$

Energy Transfer and Efficiency (A1 size)
Code: SC 87

Circuit Symbols

Commonly Used Symbols

Resistor, Battery, Switch, LED, Motor, Variable Resistor, Voltmeter, Ammeter, Fuse, Diode, LDR.

Circuit Symbols (A1 size)
Code: SC 88

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Posters – Physical Processes

Electrical Current, Potential Difference and Resistance

Electrical Charge and Current

Current is the flow of electric charge. It is measured in amperes (A). The amount of charge that flows through a wire in a given time is called the current.

Current, Potential Difference and Resistance

The current through a component depends on the potential difference across it.

Resistance

Resistance is a property of a component that opposes the flow of current. It is measured in ohms (Ω).

Ohm's Law

Potential difference = current \times resistance

$V = I \times R$

Graphs

Graphs showing the relationship between current, potential difference, and resistance.

Electrical Current, Potential Difference and Resistance (A1 size)
Code: SC 89

Electrical Circuits

Series Circuit

In a series circuit, the current is the same through all components.

Parallel Circuit

In a parallel circuit, the potential difference is the same across all components.

Current

The current through a component depends on the potential difference across it.

Resistance

Resistance is a property of a component that opposes the flow of current.

Ohm's Law

Potential difference = current \times resistance

$V = I \times R$

Electrical Circuits (A1 size)
Code: SC 90

Internal Energy

Temperature Changes and Specific Heat Capacity

Heat energy is transferred to an object, causing its temperature to rise.

Specific Latent Heat

Specific latent heat is the amount of energy required to change the state of a substance without changing its temperature.

Graphs

Graphs showing the relationship between temperature, heat energy, and specific latent heat.

Internal Energy (A1 size)
Code: SC 91

Radioactive Decay

Radiation

Alpha (α), Beta (β), and Gamma (γ) radiation are emitted from unstable nuclei.

Half-Life

Half-life is the time taken for the activity of a radioactive substance to decrease by half.

Radioactive Equations

Radioactive decay is represented by nuclear equations.

Graphs

Graphs showing the relationship between activity, time, and half-life.

Radioactive Decay (A1 size)
Code: SC 92

Forces

Scalars and Vectors

Scalars are quantities that have magnitude but no direction. Vectors are quantities that have both magnitude and direction.

Resultant Forces

Resultant force is the single force that has the same effect as the original forces.

Calculating Resultant Forces

Resultant force can be calculated using vector diagrams.

Free Body Diagrams

Free body diagrams show the forces acting on an object.

Forces (A1 size)
Code: SC 93

Speed and Acceleration

Distance and Displacement

Distance is a scalar quantity. Displacement is a vector quantity.

Speed and Velocity

Speed is the rate of change of distance. Velocity is the rate of change of displacement.

Acceleration

Acceleration is the rate of change of velocity.

Graphs

Graphs showing the relationship between distance, displacement, speed, velocity, and acceleration.

Speed and Acceleration (A1 size)
Code: SC 94

Stopping Distances & Reaction Times

Stopping Distance

Stopping distance is the distance a vehicle travels from the moment the driver decides to stop to the moment the vehicle comes to a complete stop.

Reaction Time

Reaction time is the time taken for a driver to react to a hazard.

Braking Distance

Braking distance is the distance a vehicle travels from the moment the brakes are applied to the moment the vehicle comes to a complete stop.

Graphs

Graphs showing the relationship between speed, stopping distance, and reaction time.

Stopping Distance & Reaction Times (A1 size)
Code: SC 95

Waves

Energy Transfer

Waves transfer energy from one place to another without transferring matter.

Wave Properties

Wavelength, amplitude, and frequency are properties of waves.

Wave Speed

Wave speed is the speed at which the energy of a wave travels.

Graphs

Graphs showing the relationship between wavelength, amplitude, and frequency.

Waves (A1 size)
Code: SC 96

Posters – Working Scientifically

Planning

A good plan is well designed for its purpose.

Reasons to Plan an Investigation

- Make Observations:** What structures can be seen in a cell?
- Produce a Substance:** How can a salt be made from a metal carbonate?
- Test a Hypothesis:** Is the extension of a spring proportional to its weight added?
- Explore Phenomena:** What are some patterns that occur across the world?

What to Think About When Planning

What data or observations need to be collected?

- How many measurements need to be taken?
- What range of values?
- How many repeats is enough?

What apparatus and techniques should be used?

- I will use a spring held onto a clamp stand, a 50-cm ruler and a 100-g mass.
- I will measure extension by viewing the spring at eye level and taking the reading from the bottom of the spring.

How is the apparatus used to avoid accidental measurements?

- I will attach the ruler onto the clamp stand to make sure it is measuring the length of the spring accurately.

What are the possible hazards? How can the risk be reduced?

- The clamp stand could fall over. I will attach the clamp stand to the table using a clamp and make sure it is not placed over my feet.

What are the variables?

Independent variable = mass
Dependent variable = length of extension

Variables

Investigations are often performed to identify if there are patterns or relationships between two variables. One variable is changed to see how it affects another variable.

- Independent Variable:** The independent variable is the one that is changed.
- Dependent Variable:** The dependent variable is the one that is measured for each change in the independent variable. It is what is measured and is affected during the experiment.
- Controlled Variables:** Control variables are the other variables in an investigation that should be kept the same to ensure that it is the independent variable that is affecting the dependent variable to change.

Planning (A1 size)
Code: SC 98

Evaluating Data

Think this poster is to provide the student with a reference guide to help them understand and evaluate their own experimental results.

Students record and use apparatus to determine how temperature affects the rate of reaction. They observe that the reaction is faster at higher temperatures.

This was measured by measuring how long it took for the solution to become opaque in a different temperature.

Temperature (°C)	Time taken for color to disappear (s)
20	120
30	60
40	30
50	15

Precision

Measurements are precise if they are similar and close to each other.

Accuracy

An accurate measurement is one that is close to the true value. Values are measured with uncertainty.

Repeatability

Repeatability is the extent to which repeated measurements of the same quantity under the same conditions give the same results.

Reproducibility

Reproducibility is the extent to which measurements of the same quantity under different conditions give the same results.

Graphs

Graphs are used to show the relationship between two variables. The x-axis is the independent variable and the y-axis is the dependent variable.

Line Graphs

Line graphs are used to show the relationship between two variables. The x-axis is the independent variable and the y-axis is the dependent variable.

Bar Charts

Bar charts are used to show the relationship between two variables. The x-axis is the independent variable and the y-axis is the dependent variable.

Evaluating Data (A1 size)
Code: SC 99

Presenting Data

Presenting data is the way that the student uses to present their experimental results.

Customs Data

Bar chart showing the number of tourists from different countries.

Country	Number of Tourists
USA	1200
UK	800
France	600
Germany	400
Italy	300
Spain	200
Japan	100
Australia	100
Canada	100
India	100
China	100

Line Graphs

Line graphs showing the change in temperature over time.

Bar Charts

Bar charts showing the number of people using different modes of transport.

Scatter Plots

Scatter plots showing the relationship between two variables.

Flowcharts

Flowcharts showing the steps in an experiment.

Presenting Data (A1 size)
Code: SC 100

Apparatus

A poster showing various pieces of laboratory apparatus and their uses.

Safety

Wear eye protection, lab coat, and gloves. Do not touch hot apparatus. Do not drink from any apparatus.

Experimental Apparatus

- Conical flask
- Beaker
- Test tube
- Retort stand
- Clamp
- Delivery tube
- Gas syringe
- Thermometer
- Stop watch
- Measuring cylinder
- Balance
- Crucible
- Crucible tongs
- Evaporating dish
- Spot plate
- Microscope
- Microscope slide
- Microscope cover slip
- Microscope objective lens
- Microscope eyepiece
- Microscope stage
- Microscope base
- Microscope nosepiece
- Microscope body tube
- Microscope objective lens
- Microscope eyepiece
- Microscope stage
- Microscope base
- Microscope nosepiece
- Microscope body tube

Measuring Apparatus

- Conical flask
- Beaker
- Test tube
- Retort stand
- Clamp
- Delivery tube
- Gas syringe
- Thermometer
- Stop watch
- Measuring cylinder
- Balance
- Crucible
- Crucible tongs
- Evaporating dish
- Spot plate
- Microscope
- Microscope slide
- Microscope cover slip
- Microscope objective lens
- Microscope eyepiece
- Microscope stage
- Microscope base
- Microscope nosepiece
- Microscope body tube

Apparatus (A1 size)
Code: SC 101

Physical Units

International System of Units (SI Units)

Quantity Name	Unit Name	Unit Symbol
Length	metre	m
Mass	kilogram	kg
Time	second	s
Electric current	ampere	A
Thermodynamic temperature	kelvin	K
Amount of substance	mole	mol
Luminous intensity	candela	cd

Other Units

Quantity Name	Unit Name	Unit Symbol
Temperature	degrees Celsius	°C
Energy	joule	J
Frequency	hertz	Hz
Force or weight	newton	N
Pressure	pascal	Pa
Power	watt	W
Voltage (potential difference)	volt	V
Resistance	ohm	Ω
Charge	coulomb	C
Capacitance	farad	F

SI Prefixes

These are added to unit names to produce multiples and sub-multiples, or fractions, of the original unit.

Multiples			Fractions		
Factor	Name	Symbol	Factor	Name	Symbol
10 ³	kilo	k	10 ⁻³	milli	m
10 ⁶	mega	M	10 ⁻⁶	micro	μ
10 ⁹	giga	G	10 ⁻⁹	nano	n
10 ¹²	tera	T	10 ⁻¹²	pico	p
10 ¹⁵	peta	P	10 ⁻¹⁵	femto	f
10 ¹⁸	exa	E	10 ⁻¹⁸	atto	a
10 ²¹	zetta	Z	10 ⁻²¹	zepto	z
10 ²⁴	yotta	Y	10 ⁻²⁴	yocto	y

Examples

You need to be able to convert from one unit to another.

1000 m = 1 km 1000 g = 1 kg 1000 (or 10³) m (or 1 km) 10⁻³ g (or 0.001 g) = 1 mg

Physical Units (A1 size)
Code: SC 102

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Snapframes are designed for quick and hassle free changeover of posters. They are sold fully assembled and ready to use, with clear PVC cover sheet to protect your poster. The frames are light enough to wall-mount with 4 screws.



All four sides of the Snap Frames can be easily snapped open. Just lift up four sides of frame by hand, insert your poster, place protective sheet on top and then snap frames to close without tools, as easy as 1-2-3!

Features & Benefits of Snap Frame

- Simple - access on all sides of the frame, simply flip open and change posters!
- Good visibility - clear and non-reflective PVC cover
- Durable - made of lightweight yet strong aluminium
- Instant - requires no assembly and arrives ready to use.
- Quick - Change poster from the front, no tools required!
- Eye catching - attracts attention of customers walking past
- Stable - Easy and quick wall mount with use of screws or poster hooks

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Lamination service (A1 size)	-	20.00		
Delivery(for orders < \$150)		15.00		

Please fax form to 65199196. An e-mail will be sent to confirm stock availability. Please allow 3-8 weeks for delivery.

Grand Total

Confirmed & Signed by: Name: E-mail: Contact No:

Signature & Organisation Stamp School : Address: Date: